

Name: KEY

Names: _____

Unit 5 Group Quiz – Answer Sheet

Multiple choices: (2 pts. each)

- 1) C 2) B 3) A 4) D 5) B
6) B 7) C 8) D 9) D 10) B

Chemical NAMES: (4 pts. each)

11. magnesium carbonate
12. dicarbon monoxide
13. ammonium phosphide
14. sulfur dioxide
15. calcium nitrate
16. zinc hydroxide
17. sodium bromide
18. trinitrogen pentoxide
*19. manganese (II) chloride
*20. cobalt (II) sulfide

Chemical FORMULAS: (4 pts. each)

21. $\text{Fe}^{+2} \text{Br}^{-1} = \text{FeBr}_2$
22. NO
23. P_3O_6
24. $\text{Ca}^{+2} \text{OH}^{-1} = \text{Ca}(\text{OH})_2$
25. $\text{Ag}^{+} \text{SO}_4^{-2} = \text{Ag}_2\text{SO}_4$
26. PCl_5
27. $\text{Co}^{+4} (\text{ClO}_3)^{-1} = \text{Co}(\text{ClO}_3)_4$
28. $\text{Zn}^{+2} \text{SO}_4^{-2} = \text{ZnSO}_4$
29. SiO_3
30. $\text{Cu}^{+1} \text{NO}_3^{-1} = \text{CuNO}_3$

Name:

Unit 5 Group Quiz

Choose the best answer for the following :

1. What kind of nomenclature is the following: C_2O_4
 - a. Binary Ionic
 - b. Polyatomic Ionic
 - c. Covalent
 - d. Metallic
2. What kind of nomenclature is the following: $Mn(ClO_3)_2$
 - a. Binary Ionic
 - b. Polyatomic Ionic
 - c. Covalent
 - d. Metallic
3. Covalent nomenclature is between _____.
 - a. nonmetal and nonmetal
 - b. 2 metals
 - c. metals and nonmetals
 - d. 2 cations
4. What is the name of the following monoatomic ion: P^{-3}
 - a. Phosphorous ion
 - b. Phosphorous
 - c. Phosphide (III) ion
 - d. Phosphide ion
5. What is correct name for the following: CuO
 - a. Copper oxide
 - b. Copper (II) oxide
 - c. Copper (I) oxide
 - d. Copper oxygen
6. What is formula for the following: dinitrogen tetroxide
 - a. NO_2
 - b. N_2O_4
 - c. N_4O_2
 - d. N_2O_3
7. When are Roman numerals used?
 - a. For covalent nomenclature.
 - b. For univalent ionic nomenclature.
 - c. For multivalent (transition metals).
 - d. For anions.
8. What is the name of the following monoatomic ion: N^{-3}
 - a. Nitrogen ion
 - b. Nitrogen
 - c. Nitride (III) ion
 - d. Nitride ion
9. A group of covalently bonded atoms that exhibits an overall charge is called a(n) _____.
 - a. metal
 - b. monoatomic ion
 - c. polysaccharide
 - d. polyatomic ion
10. What is formula for the following monoatomic ion: Sulfide ion
 - a. S^{-1}
 - b. S^{-2}
 - c. S^{-3}
 - d. S

Name:

Name the following: Label the following as an ionic (I) or covalent (C) first, then give the name.

___ 11. MgCO_3 _____

___ 16. Zn(OH)_2 _____

___ 12. C_2O _____

___ 17. NaBr _____

___ 13. $(\text{NH}_4)_3\text{P}$ _____

___ 18. N_3O_5 _____

___ 14. SO_2 _____

___ 19. MnCl_2 _____

___ 15. $\text{Ca(NO}_3)_2$ _____

___ 20. CoS _____

Write the chemical formula: Label the following as an ionic (I) or covalent (C) first, then give the formula.

___ 21. Iron (II) bromide _____

___ 22. Nitrogen monoxide _____

___ 23. Triphosphorus hexoxide _____

___ 24. Calcium hydroxide _____

___ 25. Silver sulfate _____

___ 26. Phosphorus pentachloride _____

___ 27. Cobalt (IV) chlorate _____

___ 28. Zinc sulfate _____

___ 29. Silicon trioxide _____

___ 30. Copper (I) nitrate _____