**Formulas for the test**

Conversion given: 1 atm = 760 mmHg = 101.3 kPa K = 0C +273

Dalton’s law of partial pressure:

Ptotal = Pa + Pb + Pc + …..

Ideal Gas Law:

PV = nRT R = 0.0821 L atm/ K mol

Combined Gas Law:

Gas Laws:

Boyle’s:

Charles’:

Gay-Lussac:

Avogadro:

Graham’s Law of Effusion: