

Electrons in the Atom – HW 4.2, 4.3, 4.4

Element	<i>Orbital Spin Notation</i>	HW 4.2	# of Valence electrons HW 4.4	Lewis Dot Diagram HW 4.4	Ion formed HW 4.4
	<i>Electron configuration</i>	HW 4.3			
	<i>Shorthand configuration</i>	HW 4.3			
1. Calcium $p^+ = \underline{\hspace{1cm}}$ $e^- = \underline{\hspace{1cm}}$					
2. Arsenic $p^+ = \underline{\hspace{1cm}}$ $e^- = \underline{\hspace{1cm}}$					
3. Cadmium $p^+ = \underline{\hspace{1cm}}$ $e^- = \underline{\hspace{1cm}}$					
4. Fluorine $p^+ = \underline{\hspace{1cm}}$ $e^- = \underline{\hspace{1cm}}$					
5. Iron $p^+ = \underline{\hspace{1cm}}$ $e^- = \underline{\hspace{1cm}}$					
6. Argon $p^+ = \underline{\hspace{1cm}}$ $e^- = \underline{\hspace{1cm}}$					
7. Sulfur $p^+ = \underline{\hspace{1cm}}$ $e^- = \underline{\hspace{1cm}}$					

<p>8. Aluminum</p> <p>$p^+ = \underline{\hspace{2cm}}$</p> <p>$e^- = \underline{\hspace{2cm}}$</p>				
<p>9. Chromium</p> <p>$p^+ = \underline{\hspace{2cm}}$</p> <p>$e^- = \underline{\hspace{2cm}}$</p>				
<p>10. Sodium</p> <p>$p^+ = \underline{\hspace{2cm}}$</p> <p>$e^- = \underline{\hspace{2cm}}$</p>				
<p>11. Oxygen</p> <p>$p^+ = \underline{\hspace{2cm}}$</p> <p>$e^- = \underline{\hspace{2cm}}$</p>				
<p>12. Sodium Ion (Na^{+1})</p> <p>$p^+ = \underline{\hspace{2cm}}$</p> <p>$e^- = \underline{\hspace{2cm}}$</p>				
<p>13. Calcium Ion (Ca^{+2})</p> <p>$p^+ = \underline{\hspace{2cm}}$</p> <p>$e^- = \underline{\hspace{2cm}}$</p>				
<p>14. Nitride Ion (N^{-3})</p> <p>$p^+ = \underline{\hspace{2cm}}$</p> <p>$e^- = \underline{\hspace{2cm}}$</p>				
<p>15. Fluoride Ion (F^{-1})</p> <p>$p^+ = \underline{\hspace{2cm}}$</p> <p>$e^- = \underline{\hspace{2cm}}$</p>				